

Mark schemes

Q1.**C**

The proportion of successful collisions increases because there is a decrease in the activation energy.

[1]**Q2.****C**

The total number of molecules in the reaction mixture

[1]**Q3.****C**

The proportion of successful collisions increases because there is a decrease in activation energy.

[1]**Q4.****C**

line C

[1]**Q5.****D**

total number of particles present

[1]**Q6.****C**

The proportion of successful collisions increases because the catalysed reaction has a lower activation energy.

[1]**Q7.****B**

At a fixed temperature their average kinetic energy is constant

[1]

Q8.**D**

There is an increase in the most probable energy of the molecules.

[1]